

Guided Notes The Atom

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Unit 7 Atomic Structure Concept 1 Notes

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Atomic Structure Concept 2 Notes

Structure of Atom Class 11 One Shot | NEET 2020 Preparation | NEET Chemistry | Arvind Arora

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Guided Notes - Teacher Edition Copyright © iTeachly.com 3 above comes after the H. This means there are two H's (hydrogen atoms) in each molecule of H₂O. 2. Coefficients - the regular-sized numbers that tell you how many molecules you have. For example, in 3H₂O, the 3 is the coefficient. This shows that there are three H₂O molecules in

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Guided Notes: The Atom History of the Atom Democritus- Greek philosopher; 400 BC 1st to use word "atomos" (means____) Aristotle-refuted what Democritus said; all matter comes from "hyle" Early Alchemists-1200-1500's; tried to turn material into ____; kept chemistry alive;

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Guided Notes: Atomic Structure 01 Atom: the smallest particle of an element that retains its ____ in a _____. Democritus believed that atoms are ____ & ____ Think about it: Was Democritus correct? Two issues with Democritus's ideas: 1. His ideas did not explain _____.

Guided Notes: Atomic Structure - KING'S SCIENCE PAGE

Guided Notes on Thomson's Model of the Atom Original Theories of Matter: 1) Democritus: matter is made up of indivisible particles called atoms 2) Dalton: elements, compounds, law of conservation of mass Which of the above theories of matter needs to be revised to become more accurate now that we know the results of Thomson's experiments with electrons?

Guided Notes on Thomsons Model of the Atom - Guided Notes ...

The Structure of the Atom Glossary TERM DEFINITION anion an ion that has a negative charge atom atomic mass atomic weight the average atomic mass of a naturally occurring sample of an element cation compound a substance made up of two or more chemical elements that are chemically bonded together electron element energy level ion isotope mass number the total number of protons and neutrons in an atom neutron

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a subatomic particle that has about the same mass as a proton but does not have an ...

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the atom remained electrically neutral (i.e. had the same number of negative and positive charges); a neutral particle would also have to be present. This led to the discovery and measurement of the mass of the neutron by Chadwick in 1932, completing the picture of the subatomic particles which makes up the atom.

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Atomic Structure Guided Notes Answers

1.____ The atomic number is equal to this particle of the atom. 2.____ This particle has no charge. 3.____ These make up the mass of an atom. 4.____ This particle has a negative charge. 5.____ These are found in the nucleus of an atom. Use the student responses as a formative assessment to determine if further instruction on atoms is warranted.

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Unit 4B- Electron Configuration- Guided Notes Atomic Structure Electrons are arranged in _____ or _____ around the nucleus of an atom o First shell can hold a maximum of _____ electrons o Second shell can hold a maximum of _____ electrons o Third shell can hold a maximum of _____ electrons ...

Name: Period: Unit 4B- Electron Configuration- Guided Notes

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